To understand how effective the Ars AB arsenate pump is. That is, how well it can remove arsenate from the cytoplasm, we need to consider the energetics.

Our starting point is to consider the arsenate efflux as a vectorial chemical reaction.

$$ATP + AsO_4^{3-}{}_{inside} \rightleftharpoons ADP + P_i + AsO_4^{3-}{}_{outside}.$$

As_i ; As_o

At equilibrium, the total Gibbs free energy is the sum of the Gibbs free energy for ATP hydrolysis and the chemical potential for "As".

$$\Delta G_{TOT} = n \cdot \Delta\mu_{As} + \Delta G_{ATP} = 0 \quad \text{(equilibrium)}$$

n: stoichiometry (= 1 in our case)

now, $\Delta G_{ATP} = \Delta G^\circ + 2.3\,RT \log_{10} \frac{[ADP][P_i]}{[ATP]}$ (products / reactant)

The standard Gibbs free energy will vary with $[Mg^{2+}]$ and pH. It is in the range of 7 - 10 kcal/mole.

$$\Delta\mu_{As} = \#\,2.3\,RT \log_{10} \frac{[As_i]}{[As_o]} + zF\Delta\Psi$$

z: valence

Since $\Delta G_{TOT} = 0$, we can equate the two (ΔG_{ATP} & $\Delta\mu_{As}$)

$$\Delta G^\circ + 2.3RT \log_{10} \frac{[ADP][P_i]}{[ATP]} = 2.3RT \log_{10} \frac{[As_i]}{[As_o]} + zF\Delta\psi$$

Now, there are real complications associated with the use of energetics. First, the reactants and products are more complex than "just" ADP, P_i, & ATP, and we are often challenged to get accurate quantitation of concentrations in the cytoplasm.
Second, we don't know the stoichiometry.
If two "As" are transported:

$$ATP + 2As_i \rightleftharpoons ADP + P_i + 2As_o$$

The equilibrium $\frac{\text{products}}{\text{reactants}}$ is

$$\frac{[ADP][P_i][As_o]^2}{[ATP][As_i]^2}$$

Finally H_3AsO_4, $H_2AsO_4^-$, $HAsO_4^{2-}$ or AsO_4^{3-} ?
Which?
It makes a big difference in the equilibrium concentrations of As_i at a given ΔG_{ATP}.

The energetics
Even so, it offers insight into the relative ability of the bacterial cell to exclude As.